Chemical Bonds

Metallic Bonding
- They are free to move
- Electrons are delocalised
- Occurs between two metals and the electrostatic force between the 'sea' of delocalised electrons and the positive nucleus of the metal
- Good conductors of electricity

Covalent Bonding
- Occurs between two non-metals
- Electrons are shared between two elements
- Strong bonds
- High melting points
- To form complete shells

Ionic Bonding
- Occurs between a nonmetal and a metal
- The metal atom loses an electron and becomes a positive ion
- The non-metal atom gains an electron and becomes a negative ion
- Electrostatic attraction occurs
- Strong bonds
- High melting points